

Which of the following statements correctly describes the action of bicarbonate, a buffer in your blood?

- A Bicarbonate releases hydrogen ions when the pH is low.
- B Bicarbonate combines with hydrogen ions when the pH is low.
- C Bicarbonate releases hydrogen ions when the blood is too acidic.
- D Bicarbonate helps to stabilize blood pH at about 5.

If the blood pH is too low, there is an excess of hydrogen ions. Bicarbonate combines with these ions, increasing the pH. Bicarbonate helps stabilize blood pH at about 7.3 to 7.5.

Answer **B**.

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